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Jennie L. Borders. Section 5.1 - Models of the Atom. The

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Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

### **Chapter 5 - Electrons in Atoms**

...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how: Aufbau principle - electrons enter the lowest energy first. This causes difficulties because of the overlap of orbitals of different energies - follow the diagram! Pauli Exclusion Principle - at most 2 electrons per orbital ...

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Chapter 5 - electrons in atoms (handouts) Chapter 6 - periodic table & trends (handouts) Chapters 7/9 - ionic bonding & naming (handouts) Chapters 8/9 - covalent bonding & chemical names & formulas (handouts) Chapters 8/15 - VSEPR/polar bonding/IMFs

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136 Chapter 5 • Electrons in Atoms Section 55.1.1 Figure 5.1  
Different elements can have similar reactions with water.  
Objectives Compare the wave and particle natures of light.  
Define a quantum of energy, and explain how it is related to an energy change of matter. Contrast continuous electromagnetic

### Chapter 5: Electrons in Atoms

Chapter 5 Electrons in Atoms. Educators. AY IB Chapter Questions. 02:25. Problem 1 Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of  $4.90 \times 10^{-7} \text{ m}$ . What is the frequency of the light? ...

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Chapter 5“Electrons in Atoms ... Describe the energies and positions of electrons according to the quantum mechanical model. Section 5.1Models of the Atom. OBJECTIVES: Describe how the shapes of orbitals related to different sublevels differ. Ernest Rutherford’s Model.

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