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covalent. 2 atoms in single covalent bond share h.... 2 atoms in double covalent bond

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share h... 2 atoms in  
triple covalent bond  
share h.... nonmetals  
bond with each other  
by sharing electrons to  
become.... 2.

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Section 16.2 - Bonding  
Theories Section

Review 16.2 13. Use  
the molecular orbital  
theory to describe  
covalent bonding.

What occurs during  
hybridization? When  
two atoms combine in  
a bonding situation,

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their atomic orbitals overlap to produce molecular orbitals. In hybridization, several atomic orbitals mix to form

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## Bonding Answer Key

Covalent Bonds 23.

Describe the difference between an ionic and a covalent bond. 24.

Classify the following compounds as ionic or covalent: a.  $MgCl_2$  b.

$Na_2S$  c.  $H_2O$  d.  $H_2S$

25. Explain why neon is monatomic but chlorine is diatomic.

26.

## **Bonding Chapter 15 16 Assignment & Problem Set**

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## Bonding Answer

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Chapter 8 Covalent

Bonding Worksheet

Answers. Classify each

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## Bonding Answer

Key  
of the following as an atom or molecule:  
a. As you should remember, ionic compounds are solids at room temperatures that have one ion strip the electron(s) from the other elements' electron. nonpolar covalent bond. atoms in a bond share the bonding electrons equally. structural formula ...

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Covalent Bonding  
Chapter 8 Answer A  
covalent bond in which  
the bonding electrons  
are most likely to be  
found in sausage-  
shaped regions above  
and below the bond  
axis of the bonded  
atoms Polar molecule A  
molecule in which one  
side of the molecule is  
slightly negative and  
the opposite side is  
slightly positive.

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## **Covalent Bonding Chapter 8 Answer Key**

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Bonding and Molecular  
Structure 8-3 There are  
two types of repulsive  
forces between the two  
atoms. First, the nuclei  
repel because they are  
both positively  
charged. Second, the  
electrons repel

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because they are both  
negatively charged.  
The attractive forces  
between the two ...

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Answer to Draw Lewis structures to illustrate the covalent bonding in the following diatomic molecules. a. Br<sub>2</sub> b. HBr c. BrF d. IBr.

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**Draw Lewis structures to illustrate the covalent bonding ...**

bond actually looks like! Chapter 9 - Covalent Bonding I. Covalent Bonding: attractive force produced as a result of shared (pgs

189electrons. -193) A.

A \_\_\_\_\_ is formed when two or more atoms bond covalently. B.

Bonds form when there is a balance of



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attractive and  
repulsive forces  
between two atoms: C.  
Single Covalent Bond &  
Lewis Structures

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